
Inorganic Chemistry

This new text for a general chemistry course should be a very teachable edition. The order of topics on matter and the kinetic theory, introducing those points of physics, which, with the first discussion on strictly chemical topics, and in essentially a descriptive form, (namely, oxygen, hydrogen, water and chlorine) leads directly and rapidly into the Periodic Classification and the Atomic and Molecular Structure. Thus, electrovalent and covalent linkages follow in natural sequence. The chemistry of carbon is introduced early in the book since those "whose major lies in some field other than chemistry usually find this more interesting than the inorganic material which it replaces. The elementary ideas of organic chemistry are no more difficult than those of inorganic chemistry." The five chapters on carbon and its compounds present an excellent summary of the fundamental simple concepts of organic chemistry and of its application in everyday life. Chapter 19 is a review and summary chapter on the material of the first half of the book and is claimed to be of great aid to the students for the semester review. Then follows series of topics covering the chemistry of the non-metals and their compounds and including other physical concepts necessary to the development of the Periodic Law. Eight chapters are devoted to the metallic elements. The last chapter is "Colloids" and is so organized that it may be introduced early in a course or eliminated entirely as the case may be. Chapter 31 is headed "Transmission Elements." Since one finds no exact definition of this term in the Chapter the conclusion could probably be reached that the term should be that more commonly used, namely, "Transition Elements." The concept of the scientific method is constantly utilized. The industrial and everyday applications of chemistry are chosen so that they may be of cultural value as well as for use in the teaching of the science. Very suitable questions are included at the end of each chapter. The text does not include any extended mathematical or theoretical discussion of certain topics such as equilibrium, but the more descriptive, qualitative discussions are apt and adequate. One finds in the word equation on page 31, "nitrate" instead of "nitrite" in the products formed, and Charles' is misspelled on page 261.

On studying the book, one feels it would be a very serviceable book for general use.—*Laurence L. Quill.*

General and Inorganic Chemistry, by Frederick C. Irwin and G. Ray Sherwood, Wayne University. x+582 pp. Philadelphia, P. Blakiston's Son and Co., Inc. 1939.